

11/pvb

SPECIFICATION

**FUEL CELL THAT CAN STABLY GENERATE ELECTRICITY
WITH EXCELLENT CHARACTERISTICS**

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Technical Field

The present invention relates to a fuel cell that uses a proton exchange membrane as an electrolyte membrane.

10 Background Art

There are several different types of fuel cells, such as phosphoric-acid fuel cells and molten carbonate fuel cells, which are classified by the type of electrolyte they use. In recent years, fuel cells that use a proton exchange membrane of positive-ion conversion type as an electrolyte membrane have actively been developed. Such proton exchange membrane fuel cells can generate electricity with high performances at relatively low operating temperatures.

A fuel cell of this type has the following basic structure.

20 A cell unit is constructed by a proton exchange membrane on one surface of which a cathode is arranged and on the other surface of which an anode is arranged. This cell unit is sandwiched between a pair of plates each having gas channels and ribs formed between adjacent gas channels. For practical

25 use, a fuel cell is composed of a number of such cell units

placed one on top of another, so that high output can be obtained.

Each of the cathode and the anode is typically composed of an electrode catalyst layer arranged on a proton exchange
5 membrane, and a gas diffusion layer arranged thereon so as to cover the electrode catalyst layer. The gas diffusion layer, as being interposed between the electrode catalyst layer and the gas channels, enables a gas flowing on the gas channels to be diffused throughout the electrode catalyst layer, and
10 also ensures conductivity between the electrode catalyst layer and the ribs on the plate.

Because ionic conductivity of the proton exchange membrane and the electrode catalyst layer is low in a dry environment, a fuel cell is typically operated with the proton
15 exchange membrane being humid by supplying a pre-humidified fuel gas and oxidizing gas. In recent years, however, demands for more compact fuel cell systems have emerged. To meet the demands, a fuel cell of different type has been developed, to which air without being humidified is supplied as an
20 oxidizing gas.

To enable such a fuel cell to generate electricity with high performances, it is preferable to uniformly humidify the entire proton exchange membrane or the entire electrode catalyst layer. This is due to the following reason. If
25 wettability of the proton exchange membrane or of the electrode

catalyst layer varies depending on its parts, such that one part is insufficiently humidified and another part is excessively humidified, the part with insufficient water content suffers from inferior ionic conductivity, and the
5 part with excessive water content suffers from inferior gas diffusibility in the electrode catalyst layer. If this happens, the fuel cell is unable to exhibit excellent electricity generation performances.

In reality, however, water content is more liable to
10 decrease due to evaporation into an oxidizing gas, in such a region of the proton exchange membrane or of the electrode catalyst layer that is in the vicinity of an oxidizing gas inlet (at the air inlet side) than in an oxidizing gas outlet side region. Therefore, the tendency is that the proton
15 exchange membrane and the electrode catalyst may often suffer from insufficient water content.

Further, water content is more liable to decrease due to evaporation into an oxidizing gas, in parts facing the oxidizing gas channels than in parts facing the ribs. Again,
20 the tendency is that the proton exchange membrane and the electrode catalyst layer may often suffer from insufficient water content.

In fuel cells that generate electricity when supplied with an oxidizing gas (air) that has not been humidified,
25 therefore, a proton exchange membrane and an electrode

layer, the porosity may be increased or the thickness may be decreased, so as to increase the gas diffusibility.

According to this technique, the oxidizing gas inlet side of the proton exchange membrane 401 can be prevented from drying and the wettability can be made uniform throughout the oxidizing gas inlet side and the oxidizing gas outlet side of the proton exchange membrane 401. However, the problem still lies in uneven wettability in the parts facing oxidizing gas channels and the parts facing ribs. This problem needs to be addressed to improve performances of fuel cells.

Also, when the technique for adjusting gas diffusibility depending on regions is employed as described above, the problem is that an oxidizing gas may not be diffused equally throughout the entire cathode. This may increase concentration polarization at the electrode reaction, causing the output to decrease.

Disclosure of the Invention

The object of the present invention is to provide a fuel cell that includes a cell unit formed by arranging an anode and a cathode on a proton exchange membrane and that can stably generate electricity with excellent characteristics, by making wettability of the proton exchange membrane and an electrode catalyst layer uniform.

To achieve the above object, in a fuel cell that includes

a cell unit formed by arranging an anode and a cathode on a proton exchange membrane, at least one of (a) a gas diffusion layer interposed between a cathode catalyst layer and a cathode side plate and (b) the cathode catalyst layer is constructed in such a manner that water retentivity is higher in its parts facing oxidant channels than in its parts facing ribs.

Here, "water retentivity of a layer" refers to an amount of water per unit volume that can be retained in the layer. The larger the amount, the higher the water retentivity of the layer.

Such water retentivity adjustment of the cathode side gas diffusion layer or the cathode catalyst layer can decrease the tendency that parts of the proton exchange membrane or the electrode catalyst layer that face the oxidant channels are more liable to dry than parts of the proton exchange membrane or the electrode catalyst layer that face the ribs. Therefore, the proton exchange membrane or the electrode catalyst layer can be uniformly humidified.

Here, the water retentivity adjustment of the gas diffusion layer is particularly effective at least in an oxidant inlet side region. This is due to the following reason. In the oxidant inlet side region, water content in an oxidizing gas is small, and therefore, water is more liable to be diffused from the gas diffusion layer into the oxidizing gas. Accordingly, wettability of the proton exchange membrane and

the electrode layer is liable to be uneven in the parts facing the oxidant channels and the parts facing the ribs.

To produce the effect of providing uniform wettability, it is also preferable to perform the above water retentivity adjustment in a predetermined range from an oxidant inlet side end toward an oxidant outlet side within the gas diffusion layer (preferably in a range of 10% to 90% of the entire region between the oxidant inlet side end and the oxidant outlet side end within the gas diffusion layer).

Such water retentivity adjustment in the gas diffusion layer can be accomplished by forming the gas diffusion layer from a conductive substrate containing a water repellent material, and by setting the water repellent material content smaller in its parts facing the oxidant channels than in its parts facing the ribs.

Alternatively, the above water retentivity adjustment can be accomplished by applying a mixture of carbon particles and a water repellent material on the gas diffusion layer so as to form a water retentivity adjustment layer, and by adjusting water retentivity of the water retentivity adjustment layer.

Such water retentivity adjustment in the water retentivity adjustment layer can be accomplished, for example, by using carbon particles with higher water retentivity (with a larger specific surface) in the parts facing the oxidant

channels than carbon particles used in the parts facing the ribs.

Also, to adjust water retentivity of the cathode side gas diffusion layer and the cathode catalyst layer as described
5 above, the gas diffusion layer and an intermediate water retentive layer that contains an ion exchanger may be interposed between the cathode catalyst layer and the second plate, the gas diffusion layer being arranged at the second plate side, and the intermediate water retentive layer being
10 arranged at the cathode catalyst layer side, and water retentivity of the intermediate water retentive layer may be adjusted to be higher in its parts facing the oxidant channels than in its parts facing the ribs.

It should be noted even though the above-described water
15 retentivity adjustment in the gas diffusion layer and the cathode catalyst layer, or the above-described provision of the intermediate water retentive layer is performed, gas diffusibility of the entire layer can be ensured and an oxidant can be diffused equally throughout electrodes.

20 According to the present invention, therefore, wettability and gas diffusibility can be ensured throughout the entire proton exchange membrane and the entire electrode catalyst layer, and therefore enables a fuel cell to generate electricity with high performances.

25 The present invention produces substantial effects,

particularly when applied to a fuel cell that generates electricity by supplying an oxidizing gas (air) without being humidified to a cathode.

5 Brief Description of the Drawings

These and other objects, advantages and features of the invention will become apparent from the following description thereof taken in conjunction with the accompanying drawings that illustrate a specific embodiment of the invention. In

10 the drawings:

FIG. 1 is an exploded view of a cell unit that constitutes a proton exchange membrane fuel cell relating to a first embodiment of the present invention;

FIG. 2 shows the construction of a fuel cell stack
15 relating to embodiments of the present invention;

FIG. 3 is a schematic cross section showing the construction of the cell unit relating to the first embodiment;

FIG. 4 is a characteristic diagram showing a cell voltage measurement value of each cell relating to practical example 1;

FIG. 5 is a cross section showing the construction of a cell unit relating to a second embodiment of the present invention;

FIG. 6 is a characteristic diagram showing a cell voltage
25 measurement value of each cell relating to practical example

water repellent resin are added. The cathode catalyst layer 22 and the anode catalyst layer 23 each are hermetically attached to a main surface of the proton exchange membrane 21 excluding its peripheral part, via hot pressing. A Pt catalyst, a Pt-Ru catalyst, or the like is used as a platinum catalyst. The Pt catalyst exhibits a favorable catalyst action when pure hydrogen is used as a fuel gas, whereas it can be poisoned and exhibits a poor catalyst action when carbon monoxide (CO) is contained in a fuel gas. On the other hand, the Pt-Ru is less liable to be poisoned by CO.

In FIG. 1, the anode catalyst layer 23 is shown by a broken line as it is provided on the lower surface of the proton exchange membrane 21.

The gas diffusion layers 24 and 25, also referred to as current collectors, are layers made of a conductive gas-permeable material. The gas diffusion layers 24 and 25 each are formed by a conductive porous material, such as carbon paper, to which a water repellent material is added to ensure the gas diffusibility. Here, carbon paper filled with a fluorocarbon resin as a water repellent material is assumed to be used, but carbon paper that has been subjected to a water repellent treatment and filled with carbon particles may instead be used.

The anode side separator plate 50 is formed by spouting and forming a mixture of a phenolic resin and carbon particles.

On a surface of the anode side separator plate 50 that faces the gas diffusion layer 25 (shown below in FIG. 1), channels 55 that allow a fuel gas (a hydrogen-rich reformed gas) to flow in direction "y" in the figure are formed, and also, ribs 56 are formed between adjacent channels 55.

The cathode side separator plate 60 is made of substantially the same material as that for the anode side separator plate 50. Though not being shown in this figure, oxidant channels 65 that allow an oxidizing gas (air) to flow in direction "y" in the figure (a large arrow "A" in the figure) and ribs 66 (see FIG. 3) are formed on the cathode side separator plate 60.

Further, holes 61 to 64, 41 to 44, 211 to 214, 31 to 34, and 51 to 54 (44, 214, 34, 54 are not shown in the figure) are formed at four corners respectively of the proton exchange membrane 21, a gasket 30, a gasket 40, the anode side separator plate 50, and the cathode side separator plate 60.

Among these, the holes 53, 33, 213, 43, and 63 form manifolds to supply a fuel gas to the channels 55 of the anode side separator plate 50, and the holes 51, 31, 211, 41, and 61 form manifolds to discharge the fuel gas. On the other hand, the holes 54, 34, 214, 44, and 64 are manifolds to supply an oxidizing gas to the channels of the cathode side separator plate 60, and the holes 52, 32, 212, 42, and 62 are manifolds to discharge the oxidizing gas.

For practical use, a fuel cell stack is formed to enable high voltage and high power to be obtained as shown in FIG.

2. In detail, a number of cell units 10 with the construction described above are stacked up, and these cell units 10 are supported between collector plates 70, insulator plates 71 for insulating electricity and heat, and fastening plates 72 for applying load to hold this stack structure. This stack structure is then clamped with bolts 73 and nuts 74. Note that the clamping load is applied by disc springs 75.

10 When this fuel cell stack is operated, one or both of a fuel gas and air is humidified and then supplied to respective channels, and the fuel gas and the air are used to generate electricity in the following way.

The fuel gas supplied to the channels 55 is supplied to the anode catalyst layer 23 via the gas diffusion layer 25. Within the anode catalyst layer 23, hydrogen in the fuel gas is then ionized to release electrons ($H^2 \rightarrow 2H^+ + 2e^-$). Protons generated here move toward the cathode catalyst layer 22 side within the proton exchange membrane 21. Along with this movement of protons, water content in the anode catalyst layer 23 also moves toward the cathode catalyst layer 23 side.

On the other hand, air is supplied to the cathode catalyst layer 22 side via the gas diffusion layer 24. Within the gas diffusion layer 24, oxygen in the air is bonded to the protons moving thereto within the proton exchange membrane 21, so

10 The following describes in detail the construction and
the effects of the membrane/electrode assembly 20.

The gas diffusion layer 24, sandwiched between the cathode catalyst layer 22 and the cathode side separator plate 60, is constructed in such a manner that water retentivity is adjusted in a predetermined region (20A in the figure) from the inlet side (the hole 64 side) of an oxidizing gas (air) in the following way. In the predetermined region, water retentivity is adjusted to be higher in parts facing the oxidant channels 65 (referred to as "channel facing parts 24A") than in parts facing the ribs 66 (referred to as "rib facing parts 24B").

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Here, it is preferable to set the ratio (X_a/X_b) of the water repellent material content (X_a) per unit area in the channel facing parts 24A with respect to the water repellent material content (X_b) per unit area in the rib facing parts 24B in a range of 0.2 to 0.9 inclusive. The reason for this is described later in practical example 1.

FIG. 3 is a schematic cross section of the cell unit 10 in the air inlet side region 20A, taken in the thickness direction thereof (x-z plane in FIG. 1).

By adjusting the water retentivity in the gas diffusion layer 24 as described above, the following effects can be produced.

Because air flows along the surface of the channel facing parts 24A, water content in the gas diffusion layer 24 is more liable to decrease due to evaporation into the air in the channel facing parts 24A than in the rib facing parts 24B, as indicated by a graph showing water evaporation degree in FIG. 3. In particular, air flowing through the oxidant channels 65 is drier in the air inlet side region 20A than in the air outlet side region 20B. Therefore, more water content tends to be lost in the channel facing parts 24A due to evaporation. Note that this tendency is remarkable particularly when air without being humidified is supplied to the oxidant channels 65.

25 Taking this into consideration, the present cell unit

10 Accordingly, the tendency that the parts of the proton
exchange membrane and the cathode catalyst layer facing the
oxidant channels 65 are more liable to dry can be decreased.
This allows the wettability of the parts facing the oxidant
channels 65 and the parts facing the ribs 66 in the
15 membrane/electrode assembly 20 to be made uniform.

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Taking this into consideration to provide uniform wettability throughout the gas diffusion layer 24, it is preferable to adjust the water retentivity only in the air inlet side region 20A as described above. Also, it is preferable to set the size of the air inlet side region 20A where water retentivity is to be adjusted, in a range of 10 to 90% inclusive of the size of the entire gas diffusion layer 24.

However, there are cases where air flowing through the oxidant channels 65 is extremely dry even in the air outlet vicinity, for example, when air without being humidified is supplied. In such a case, it is considered preferable to adjust the water retentivity throughout the entire gas diffusion layer 24.

Also, the water retentivity being adjusted depending on regions of the gas diffusion layer 24 does not have much effect on its gas diffusibility. Therefore, the gas diffusibility can be maintained uniform in the entire reaction region of the gas diffusion layer 24. Accordingly, electricity can be stably generated with an excellent cell voltage.

[Practical Example 1]

Based upon the above first embodiment, cells (cell units) 1 to 5 of practical examples, and cells 6 and 7 of comparative

examples were produced according to the following specification, and the cell voltage measuring experiment was carried out on each cell.

5 Table 1

	PTFE Content in Gas Diffusion Layer		PTFE Content Ratio Xa/Xb
	Channel Facing Part	Rib Facing Part	
	Xa	Xb	
Cell 1	4	30	0.13
Cell 2	6	30	0.20
Cell 3	10	30	0.33
Cell 4	17	30	0.57
Cell 5	25	30	0.83
Cell 6	30	30	1.00
Cell 7	40	30	1.33

Proton exchange membrane:

Perfluoro carbon sulfonic acid film (*Nafion*: 112 films,
size: 12*12cm, thickness: 50 μ m)

10 Cathode catalyst layer 22 and anode catalyst layer 23:
size: 10*10cm, thickness: 20 μ m

Gas diffusion layer (cathode and anode):

Carbon paper was filled with polytetrafluoroethylene
(PTFE) as a water repellent material, so as to have a size
15 of 10*10cm and a thickness of 200 μ m.

For cells 1 to 5, in the air inlet side region of the
cathode side gas diffusion layer (50% of the entire region),

5 In the air outlet side region in the cathode side gas
diffusion layer, the PTFE content was uniformly set at 30wt%.

For each cell, the PTFE content in the anode side gas diffusion layer was uniformly set at 30wt% throughout its entire region.

Cells 6 and 7 of comparative examples were produced in the same manner as cells 1 to 5 of practical examples except that the fluorocarbon fiber content in the channel facing parts 24A of the cathode side gas diffusion layer of cells 6 and 7 was respectively set at 30 and 40wt%.

15 The following describes a method for producing each cell
in detail.

Cathode side gas diffusion layer: carbon paper (thickness: 200 μ m) was prepared, immersed in a PTFE dispersion solution, and baked for one hour at 380°C. The PTFE content at this point was adjusted to be a predetermined value (4wt% for cell 1, 6wt% for cell 2, 10wt% for cell 3, 17wt% for cell 4, and 25wt% for cell 5) uniformly in the entire region.

A masking sheet was placed only on parts corresponding to the channel facing parts 24A in the air inlet side region of the layer. A PTFE dispersion solution was sprayed and

Cathode catalyst layer and anode catalyst layer: Pt support carbon particles, a *Nafion* solution, and a PTFE dispersion solution were mixed at a weight ratio of Pt support carbon particles:*Nafion*:PTFE as 100:20:10, to form a slurry.

5 This slurry was applied on each of the above formed gas diffusion layers so as to have a thickness of 20 μm , to form a cathode catalyst layer and an anode catalyst layer.

The gas diffusion layers (anode and cathode) on each of which a catalyst layer is formed were arranged one on top of the other with a proton exchange membrane being interposed
10 therebetween, and subjected to hot pressing for 60sec. at 150°C, to complete each of cells 1 to 7.

(Cell Voltage Measuring Experiment)

Cells 1 to 7 of practical examples and comparative
15 examples produced as described above were operated under the following conditions, and the cell voltage (a value of stable voltage) was measured.

Current Density: 0.5A/cm²

Operating Temperature within Cell: 80°C

20 Fuel Gas: pure hydrogen

Oxidizing Gas: air

Fuel Gas Utilization Ratio: 70%

Oxidizing Gas Utilization Ratio: 40%

FIG. 4 shows the cell voltage measurement value of each
25 cell, and is a characteristic diagram showing the relationship

between (a) the ratio (X_a/X_b) of the PTFE content X_a in the channel facing parts 24A with respect to the PTFE content X_b in the rib facing parts 24B in the air inlet side region and (b) the cell voltage (mV).

5 Considerations:

As can be seen from FIG. 4, cells 2 to 5 can produce a higher cell voltage than cells 6 and 7. Therefore, it is preferable to set the ratio (X_a/X_b) in a range of 0.2 to 0.8.

The reason for this can be considered as follows. For
 10 cells 2 to 5, the ratio (X_a/X_b) in the cathode side diffusion layer is set in a range of 0.2 to 0.8. This means that the water retentivity of the channel facing parts 24A is appropriately higher than that of the rib facing parts 24B. Therefore, the wettability of the proton exchange membrane
 15 and the electrode layer is made uniform in the channel facing parts 24A and the rib facing parts 24B. For cells 6 and 7, however, the effect of such uniform wettability is not provided.

Also, cell 1 produces a considerably low cell voltage,
 20 compared with cells 2 to 5.

The reason for this can be considered as follows. For cell 1, the fluorocarbon resin content in the channel facing parts 24 in the air inlet side region is considerably low. This allows water to reside in these parts, extremely degrading
 25 the gas diffusibility.

a conductive porous material such as carbon paper to which a water repellent material is added to ensure gas diffusibility. Besides this, a mixture of carbon particles and a water repellent material is applied on the surface of the cathode side gas diffusion layer 124 facing the cathode catalyst layer 22, so as to form a water retentivity adjustment layer 125.

The water retentivity adjustment layer 125 is superior in water retentivity as it contains carbon particles. The water retentivity adjustment layer 125 present adjacent to the cathode catalyst layer 22 enables the proton exchange membrane 21 and the cathode catalyst layer 22 to be maintained humid with stability.

In the present embodiment, the water retentivity adjustment is performed in the air inlet side region 20A of the water retentivity adjustment layer 125. To be more specific, in the air inlet side region 20A of the water retentivity adjustment layer 125, water retentivity of parts facing the oxidant channels 65 (channel facing parts 125A) is adjusted to be higher than water retentivity of parts facing the ribs 66 (rib facing parts 125B).

Such water retentivity adjustment in the water retentivity adjustment layer 125 is specifically accomplished, by using carbon particles with a large specific surface for the channel facing parts 125A and carbon particles with a small specific surface for the rib facing parts 125B.

examples were produced according to the same specification as employed in the practical experiment 1, except the cathode side gas diffusion layer. The cell voltage measuring experiment was carried out on each cell.

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Table 2

	Specific Surface of Carbon (m ² /g)		Carbon Specific Surface Ratio Ya/Yb
	Channel Facing Part Ya	Rib Facing Part Yb	
Cell 8	1270	254	5.00
Cell 9	800	254	3.15
Cell 10	343	254	1.35
Cell 11	254	254	1.00
Cell 12	56	254	0.22

For cells 8 to 10, the cathode side gas diffusion layer was formed in the following way. Carbon paper was filled with PTFE as a water repellent material (the PTFE content uniformly set at 30wt%), so as to have a size of 10*10cm and a thickness of 200 μ m. On the surface of this layer, a mixture of carbon particles and PTFE was applied, to form the water retentivity adjustment layer. In the air inlet side region of the water retentivity adjustment layer (50% of the entire region), water retentivity of the channel facing parts 125A was adjusted to be higher than that of the rib facing parts 125B, by using carbon particles with a larger specific surface for the channel facing parts 125A as shown in Table 2.

For each cell, the PTFE content in the mixture forming the water retentivity adjustment layer was set uniformly at 30wt%.

As Table 2 shows, for cell 11 of comparative example, carbon particles with the same specific surface were used for the channel facing parts 125A and the rib facing parts 125B in the water retentivity adjustment layer. For cell 12 of comparative example, carbon particles with a smaller specific surface than that for the rib facing parts 125B were used for the channel facing parts 125A in the water retentivity adjustment layer 125, to decrease the water retentivity thereof. Except for these points, cells 11 and 12 of comparative examples were produced in the same manner as the above cells 8 to 10.

Cathode side gas diffusion layer: carbon paper (thickness: 200 μ m) was prepared, immersed in a PTFE dispersion solution, and baked for one hour at 380°C. The PTFE content at this point was set uniformly in the entire region at 30wt%.

The mixture dispersion solution was applied thereon.

Following this, five types of carbon particles respectively having specific surface values of 1270, 800, 343, 254, and $56\text{m}^2/\text{g}$ each were mixed with the PTFE dispersion solution to form five mixture slurries (the PTFE content within the solid content: 30wt%). A masking sheet having apertures corresponding to the channel facing parts 125A in the air inlet side region was placed on the layer. The corresponding one of the five mixture slurries was applied thereon. Then, the layer was baked for one hour at 380°C .

As described above, the cathode side gas diffusion layers for cells 8 to 12 were completed. Using these cathode side gas diffusion layers, cells 8 to 12 were produced as in the above practical example 1.

15 (Cell Voltage Measuring Experiment)

The cells of practical examples and comparative examples produced as described above were operated under the same conditions as employed in the cell voltage measuring experiment in the practical example 1, and the cell voltage

20 was measured.

FIG. 6 shows the cell voltage measurement value of each cell, and is a characteristic diagram showing the relationship between (a) the ratio (Y_a/Y_b) of a specific surface Y_b of a carbon particle for use in the rib facing parts 125A with respect to a specific surface Y_a of a carbon particle for

Considerations:

As can be seen from FIG. 6, cells 8 to 10 can produce
5 a higher cell voltage than cells 11 and 12.

The reason for this can be considered as follows. For cells 8 to 10, the ratio (Y_a/Y_b) of the specific surface in the air inlet side region of the water retentivity adjustment layer 125 was set in a range of 1.35 to 5. Therefore, water retentivity of the channel facing parts 125A is appropriately higher than that of the rib facing parts 125B. Due to this, for cells 8 to 10, wettability of the proton exchange membrane and the electrode layer is made uniform in the channel facing parts 125A and the rib facing parts 125B, whereas for cells 11 and 12, the effect of such uniform wettability is not provided.

[Third Embodiment]

A cell unit relating to the present embodiment has the same construction as the cell unit 110 in the second embodiment, with the only difference being in a method for adjusting water retentivity in the water retentivity adjustment layer 125.

To be more specific, in the second embodiment, the water retentivity of the channel facing parts 124A and the rib facing parts 125B in the water retentivity adjustment layer 125 is

retentivity thereof. Except for these points, cells 18 and 19 of comparative examples were produced in the same manner as the above cells 13 to 17.

The following describes a method for producing each cell
5 in detail.

Cathode side gas diffusion layer: carbon paper (thickness: 200 μm) was prepared, immersed in a PTFE dispersion solution, and baked for one hour at 380°C. The PTFE content at this point was set uniformly in the entire region at 30wt%.

Carbon particles with a specific surface of $254\text{m}^2/\text{g}$ and a PTFE dispersion solution were mixed to form a mixture slurry (the PTFE content within the solid content: 30wt%). Then, a masking sheet was placed only on parts corresponding to the channel facing parts 124A in the air inlet side region. The mixture dispersion solution was applied thereon.

Following this, carbon particles with a specific surface of 254m²/g and the PTFE dispersion solution were mixed to form mixture slurries whose PTFE content within the solid content is respectively set at 4, 6, 10, 17, 25, 30, and 40wt%. Then, a masking sheet having apertures corresponding to the channel facing parts 125A in the air inlet side region was placed on the layer. The corresponding one of the mixture slurries was applied thereon, and the layer was baked for one hour at 380°C.

25 As described above, the cathode side gas diffusion layers

for cells 13 to 19 were completed. Using these cathode side gas diffusion layers, cells 8 to 12 were produced in the same manner as in the practical example 1.

(Cell Voltage Measuring Experiment)

5 The cells of practical examples and comparative examples produced as described above were operated under the same conditions as employed in the cell voltage measuring experiment in the practical example 1, and the cell voltage was measured.

10 FIG. 7 shows the cell voltage measurement value of each cell, and is a characteristic diagram showing the relationship between (a) the ratio (Z_a/Z_b) of the PTFE content (Z_b) in the rib facing parts 125 with respect to the PTFE content (Z_a) in the channel facing parts 125B in the air inlet side
15 region of the water retentivity adjustment layer 125 and (b) the cell voltage (mV).

Considerations:

As can be seen from FIG. 7, cells 14 to 17 can produce a higher cell voltage than cells 18 and 19. From this, it
20 is considered preferable to set, within the water retentivity adjustment layer 125, the ratio of the fluorocarbon resin content in the channel facing parts 125A with respect to that in the rib facing parts 125B, in a range of 0.2 to 0.8.

The reason for this can be considered as follows. For
25 cells 14 to 17, in the air inlet side region of the water

In the present embodiment, the cathode side gas diffusion layer 24 is made of a conductive porous material such as carbon paper to which a water repellent material is uniformly added to ensure the gas diffusibility.

Further, an ion exchanger resin is applied on the surface of the cathode catalyst layer 22 on the cathode side gas diffusion layer 24, to form the intermediate water retentive layer 26 therebetween.

The intermediate water retentive layer 26 is superior in water retentivity as it is made of the ion exchanger resin. The intermediate water retentive layer 26 present adjacent to the cathode catalyst layer 22 enables the proton exchange membrane 21 and the cathode catalyst layer 22 to be maintained humid with stability.

Also, the water retentivity of the intermediate water retentive layer 26 is adjusted to be higher in its parts facing the oxidant channel 65 (channel facing parts 26A) than in its parts facing the ribs 66 (rib facing parts 26B).

Such water retentivity adjustment in the intermediate

water retentive layer 26 is specifically accomplished, by increasing an applied amount of the ion exchanger resin in the channel facing parts 26A and decreasing an applied amount of the ion exchanger resin in the rib facing parts 26B (or
5 not applying the ion exchanger resin in the rib facing parts 26B).

Alternatively, the water retentivity adjustment can be accomplished by using an ion exchanger resin with a larger ion-exchange capacity in the channel facing parts 26A and
10 using an ion exchanger resin with a smaller ion-exchange capacity in the rib facing parts 26B.

By providing the intermediate water retentive layer 26 whose water retentivity is adjusted in the above described way between the cathode catalyst layer 22 and the cathode
15 side gas diffusion layer 24, the same effects as described in the first embodiment can be produced. To be more specific, the wettability of the proton exchange membrane 21 and the cathode catalyst layer 22 can be made uniform in the parts facing the oxidant channels 65 and the parts facing the ribs
20 66.

Also, it is preferable to set the size of the air inlet side region 20A where the intermediate water retentive layer 26 is to be provided, in a range of 10 to 90% inclusive of the size of the entire gas diffusion layer 24 for the same
25 reason described in the first embodiment.

retentive layer 26 formed on the cathode side gas diffusion layer and not on the rib facing parts 26B of the intermediate water retentive layer 26. Accordingly, in the intermediate water retentive layer 26, water retentivity of the channel facing parts 26A is higher than that of the rib facing parts 26B.

For cell 21, in the air inlet side region on the surface of the cathode side gas diffusion layer (50% of the entire region), an ion exchanger resin with an ion-exchange capacity of 1.0meq./g was applied, by an amount of 0.06mg/cm², on the parts corresponding to the channel facing parts 26A in the same manner as that for cell 20. Then, in the air inlet side region on the surface of the cathode side gas diffusion layer (50% of the entire region), a masking sheet having apertures corresponding to the rib facing parts 26B was placed, an ion exchanger resin (an alcohol solution) with an ion-exchange capacity of 0.91meq./g was applied thereon, and was dried for 20 min. at 80°C to remove alcohol. The applied amount of the ion exchanger resin in the rib facing parts 26B was also set at 0.06mg/cm².

In this way, the ion exchanger resin with a relatively high ion-exchange capacity was applied only on the channel facing parts 26A of the intermediate water retentive layer 26 formed on the cathode side gas diffusion layer, and the ion exchanger resin with a relatively low ion-exchange

capacity was applied on the rib facing parts 26B of the intermediate water retentive layer 26. Accordingly, in the intermediate water retentive layer 26, water retentivity of the channel facing parts 26A can be made higher than that of the rib facing parts 26B.

(Comparative Example)

Also, as comparative examples, cells 22 to 29 were produced in the same manner as that for the above cells 20 and 21, except that an ion exchanger resin with an ion-exchange capacity of 1.0meq./g was uniformly applied in the air inlet side region (50% of the entire region) on the surface of the cathode side gas diffusion layer, to form an intermediate water retentive layer.

Here, for cells 22 to 29, the applied amount of the ion exchanger resin was set at values shown in Table 4.

Table 4

	Ion Exchanger Resin Applied Amount • Ion-Exchange Capacity	
	Channel Facing Part	Rib Facing Part
Cell 20	0.06mg/cm ² 1.0meq./g	0
Cell 21	0.06mg/cm ² 1.0meq./g	0.06mg/cm ² 0.91meq./g
Cell 22	0.01mg/cm ²	1.0meq./g
Cell 23	0.02mg/cm ²	1.0meq./g
Cell 24	0.04mg/cm ²	1.0meq./g

Note that cell 30 was produced in the same manner as that for cells 20 and 21 except that an ion exchanger resin was not applied (an intermediate water retentive layer was not formed).

(Cell Voltage Measuring Experiment)

The cells of practical examples and comparative examples produced as described above were operated under the following conditions, and the cell voltage was measured.

10 Current Density: $0.5\text{A}/\text{cm}^2$

Operating Temperature within Cell: 80°C

Fuel Gas: pure hydrogen

Oxidizing Gas: air

Fuel Gas Utilization Ratio: 70%

15 Oxidizing Gas Utilization Ratio: 40%

Gas Humidifying Temperature: fuel gas being humidified
at 80°C, and oxidizing gas not being humidified

The experimental results showed that the cell voltage measurement value was 632mV for cell 20 of practical example, and 643mV for cell 21 of practical example.

For cell 25 of comparative example, the cell voltage measurement value was 634mV.

FIG. 9 shows the cell voltage measurement value of each cell.

5 Considerations:

As can be seen from FIG. 9, cells 23 to 28 with the applied amount of the ion exchanger resin being in a range of 0.02 to 0.12mg/cm² can produce a higher cell voltage than cells 23 and 30 with the applied amount of the ion exchanger resin being 0.01mg/cm² or less and cell 29 with the applied amount of the ion exchanger resin being 0.15mg/cm².

These experimental results indicate the following. When the cell is operated with an oxidizing gas not being humidified, a large amount of water content is lost due to evaporation from the membrane/electrode assembly if the intermediate water retentive layer is not provided. On the other hand, an intermediate water retentive layer with appropriate water retentivity, if provided, can prevent the water evaporation from the membrane/electrode assembly, thereby enabling a higher cell voltage to be obtained.

The experimental results also indicate that it is difficult to obtain a high cell voltage when an excessively large amount of ion exchanger resin is applied to form the intermediate water retentive layer. This can be considered due to the following reason. Because the ion exchanger resin

itself is non-conductive, too much ion exchanger resin being applied causes the electrical resistance between the cathode catalyst layer and the gas diffusion layer to increase to a great degree.

5 Next, comparing cell 20 of practical example with cell
25 of comparative example, the total applied amount of ion
exchanger resin of cell 20 is about half of that of cell 25.
However, cells 20 and 25 each produce a similar cell voltage.
Also, comparing cell 20 of practical example with cell 23
10 or with cell 24 of comparative example, cell 20 of practical
example produces a higher cell voltage than cells 23 or 24
although the total applied amount of ion exchanger resin of
these cells is the same.

The reason for this can be considered as follows. For cell 24 of practical example, in the air inlet side region, the water retentivity of the channel facing parts is appropriately higher than that of the rib facing parts in the intermediate water retentive layer. Therefore, the wettability of the proton exchange membrane and the electrode layer is made uniform in the channel facing parts and the rib facing parts, whereas for the cells of comparative examples, the effect of such uniform wettability is not provided.

[Fifth Embodiment]

25 FIG. 10 is a schematic cross section of a cell unit

relating to the present embodiment in the air inlet side region,
taken in the thickness direction thereof.

A cell unit 310 in the present embodiment has the same construction as the cell unit 10 shown in FIG. 1 in the first embodiment, except that the water retentivity is not adjusted in a cathode side gas diffusion layer 324 but is adjusted in a cathode catalyst layer 322 in a membrane/electrode assembly 220.

In the present embodiment, the cathode catalyst layer 322 is made of platinum catalyst support carbon as the above cathode catalyst layer 22. However, the water retentivity of channel facing parts 322A facing the oxidant channels 65 is set higher than that of rib facing parts 332B facing the ribs 66 in the present embodiment.

15 This water retentivity adjustment in the cathode catalyst layer 322 is specifically accomplished, by using carbon particles with a larger specific surface (carbon particles with a high water retentivity) as a catalyst support in the channel facing parts 322A and using carbon particles
20 with a smaller specific surface (carbon particles with a low water retentivity) in the rib facing parts 222B.

By providing such a cathode catalyst layer 322 whose water retentivity has been adjusted, the same effects as described in the first embodiment can be produced.

25 To be more specific, the wettability of the proton

combined for use.

The second embodiment describes the case where the water retentivity adjustment layer made of a mixture of carbon particles and a water repellent material is formed only on the cathode side gas diffusion layer. However, another water retentive adjustment layer made of a mixture of carbon particles and a water repellent material may be formed on the anode side gas diffusion layer. By doing so, a greater effect of humidifying the proton exchange membrane can be expected.

The above embodiments are described based on a cell unit in which an oxidizing gas and a fuel gas flow in parallel. However, the construction of the cell unit should not be limited to such. For example, the cell unit may have the construction in which channels of the cathode side separator plate and those of the anode side separator plate are perpendicular with each other.

Besides the water retentivity adjustment between the channel facing parts and the rib facing parts described above, the water repellent material content may also be adjusted in such a manner that the water repellency in the gas diffusion layer is higher in the air inlet side region 20A than in the air outlet region 20B.

By combining these adjustments between the air inlet side region and the air outlet side region, water evaporation

is suppressed in the air inlet side region 20A, as compared in the air outlet side region 20B. Therefore, the wettability of the proton exchange membrane and the catalyst layer can be made uniform in the air inlet side and the air outlet side.

5 The above embodiments are described taking a typical proton exchange membrane fuel cell that generates electricity using a fuel gas as example. However, the present invention is applicable to a direct methanol fuel cell (DMFC) that generates electricity while supplying methanol instead of
10 a fuel gas directly to the anode side.

Industrial Application

The present invention can be utilized in generators, generation systems, cogeneration systems, electric cars, and
15 the like that use a fuel cell that includes a cell unit formed by arranging an anode and a cathode on a proton exchange membrane.